



ST. JOSEPH'S COLLEGE, PRA YAGRAJ

FIRST UNIT TEST- MAY 2024

CLASS – IX

SUBJECT- CHEMISTRY

Time: 1½ Hrs

Max. Marks: 30

ALL THE QUESTIONS ARE COMPULSORY

Q 1. Choose one correct answer to the questions from the given options:

(Do not copy the question, write the correct answers only) [2]

- (a) In $(\text{NH}_4)_x \text{SO}_4$, the value of x is-
a. 2 b. 1 c. 3 d. 4
- (b) The symbol of cobalt is given as-
a. Co b. CO c. Cb d. Cd
- (c) *Kalium* is the Latin name of -
a. Krypton b. Potassium c. Lead d. Tin
- (d) In CO_2 , the carbon to oxygen mass ratio is-
a. 3:7 b. 3:5 c. 3:4 d. 3:8

Q 2. Fill in the blanks- [2]

- (a) The acidic radical in calcium oxide is _____.
- (b) The chemical formula of limestone is _____.
- (c) The chemical name of the compound CCl_4 is _____.
- (d) The higher valency of iron is _____.

Q 3. Differentiate between the following- [2]

- (a) atomic mass unit and relative atomic mass .
- (b) H_2 and 2H (where H is the symbol of hydrogen).

Q 4. Give reason: [2]

- (a) Why certain elements exhibit more than one valency?
- (b) Why should an equation be balanced?

Q 5. Mention the valency of sulphur in each of the following compound- [2]

- (a) H_2S (b) SO_2 (c) SO_3 (d) CS_2

Q 6. Define the following terms- [2]

- (a) Symbol (b) Valency (c) Molecular formula (d) Radical

Q 7. Write the chemical formula for each of the following - [2]

- (a) Aluminium Carbide (b) Sodium bicarbonate
- (c) Calcium phosphate (d) Ferrous sulphate



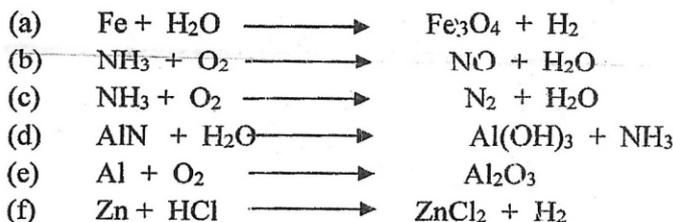
Q 8. The formula of the chloride of a metal 'M' is MCl_2 . State the formula of its - [2]
(a) Hydroxide (b) Carbonate
(c) Nitrate (d) Phosphide

Q 9. Write the chemical names of the following compounds- [2]
(a) $NaClO_3$ (b) Ca_3N_2
(c) $KMnO_4$ (d) K_2O

Q 10. Find the percentage composition of each element in - [3]
(a) CH_3COOH (b) $K_2Cr_2O_7$
[Given Relative atomic mass: C=12, H=1, O=16, K=39, Cr=52]

Q 11. Calculate the relative molecular masses of the following compounds- [3]
(a) Copper sulphate pentahydrate
(b) Sodium thiosulphate
(c) Glucose
[Given Relative atomic masses : Cu=64, S=32, O=16, H=1, Na=23, C=12]

Q 12. Balance the following chemical equations- [3]



Q13. Elements A, B and C have 2, 7 and 5 electrons in their outermost shells respectively. Write - [3]

- (a) The element(s) which is/are metal.
(b) The element(s) which gain electron(s) to become stable.
(c) The element(s) which can form trivalent ion.
(d) The formula of the compound formed between
i. A and B
ii. A and C

*****ALL THE BEST*****